

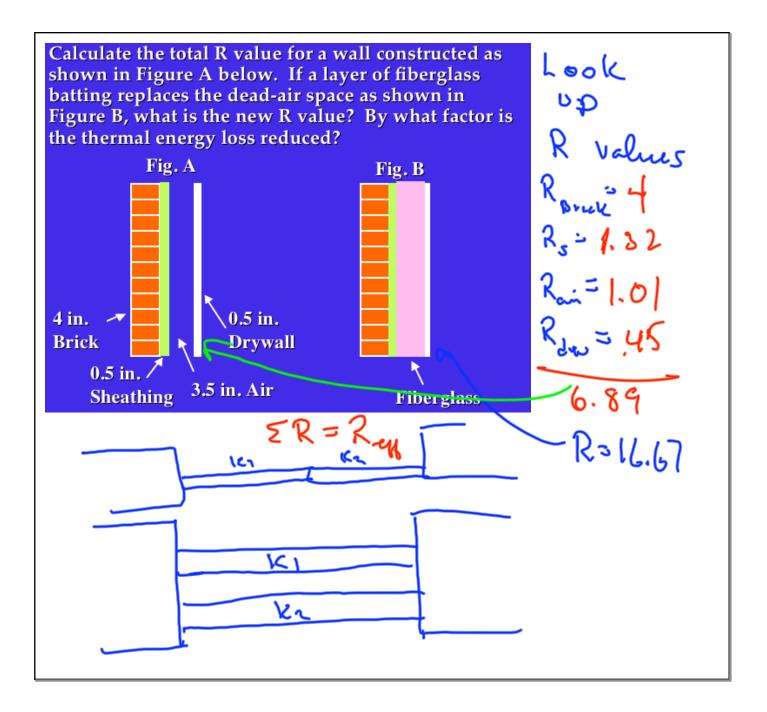
$$H_{1} = H_{2}$$

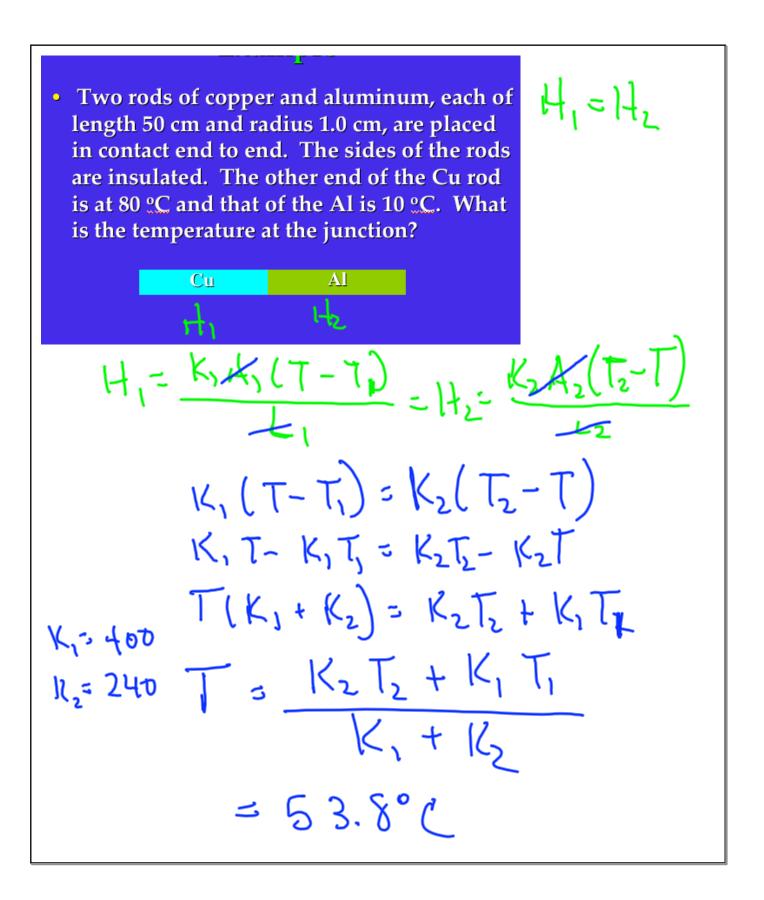
$$H_{1} = K_{1} \underline{A} \underline{A} \underline{T}_{1} \qquad H_{2} = K_{2} \underline{A} \underline{A} \underline{T}_{2}$$

$$\Delta T = \Delta T_{1} + \Delta T_{2}$$

$$\Delta T = H_{1} \underline{L}_{1} + \frac{H_{2} L_{2}}{L_{2} \underline{A}} = H \begin{bmatrix} L_{1} + L_{2} \end{bmatrix} \frac{1}{L_{1}}$$

$$H = \frac{A \Delta T}{\frac{1}{L_{1}} + \frac{1}{L_{2}}} = \frac{A \Delta T}{\frac{1}{L_{1}} + \frac{1}{L_{2}}}$$





Van der Waals  $+\left(\frac{\alpha n^2}{\sqrt{r^2}}\right)\left(\sqrt{-nb}\right)$ RT <sup>b</sup> Inter molecula Volume 8 I mole forus of molecules

An automobile tire is pumped to a gauge pressure of 200kPa when it is at 20°C. After the car have been driven at high speed, the temperature has been increased to 50°C. Assuming the volume is unchanged, find the new gauge pressure. Repeat, if the tire expands by 10 percent.

$$PV = nRT \qquad \frac{P}{T} = \frac{nR}{V} = constant$$

$$\frac{P_o}{T_o} = \frac{P_c}{T_c} \qquad P_c = \frac{T_c}{T_o} P_o$$

$$= \frac{323k}{393} (200 kR_a + 16 kR_a)$$

$$P_c = 332 kP_a$$

$$P_c = 332 kP_a - 10 (kP_s)$$

$$= 23 (kP_a)$$